

Problematic text in the manuscript	Source found in the internet	Suggested action
The prepared solution poured into the 4 glass beakers (1000 ml for each one). Then these beakers were placed on a heater at 50 ° C for 1 hour under a pressure of 0.9 bars. The solution was saturated with carbon dioxide. After that, the electrodes were placed in the medium and are connected through a potentiometer ACM GILL AC. The surface of working electrode is cleaned by acetone before using, these electrodes are using for one time. After 1 hour, except for 1 beaker, the remaining 3 is fed with the suitable amount of inhibitor and continued supply of CO₂ under pressure of 0.9 bar until the end of the experiment. The potential of the working electrode was varied by a CoreRunning programme (Version 5.1.3.) through an ACM instrument Gill AC. The CoreRunning programme converts a corrosion current in mA/cm² to a corrosion rate in mm/year. A cylindrical carbon steel rod of the composition 080A15	http://pubs.sciepub.com/jmpc/1/2/2	Kindly take the reference of the source inside your paper to avoid plagiarism complain.

GRADE STEEL was used as a working electrode. Gill AC technology allows measure DC and AC signals using standard Sequencer software. A small sweep from typically -10 mV to $+10$ mV at 10 mV/min around the rest potential is performed.		
shows that, the change in corrosion rate (CR) with time for carbon steel in CO₂-saturated 1%NaCl solution containing different concentrations form inhibitors (a) II, (b) VI and (c) V at 50 °C. The inhibitor was added after 1 hour of exposure because at this time the corrosion potential got stable, allowing the measurement of the CR prior the injection of the inhibitor. The initial corrosion rate, without inhibitor, was measured to be between 3.45 and 5.03 mm y⁻¹, but in the presence 100 ppm of the investigated inhibitors between 0.154 and 0.031 mm y⁻¹. It can be observed from Figure 1 that, the CR, in the absence of inhibitor, tends to increase with time. The increase in CR has been attributed to the	http://pubs.sciepub.com/jmpc/1/2/2	Kindly take the reference of the source inside your paper to avoid plagiarism complain.

galvanic effect between the ferrite phase and cementite (Fe_3C) which is a part of the original steel in the non-oxidized state and accumulates on the surface after the preferential dissolution of ferrite ($\alpha\text{-Fe}$) into Fe^{2+} [20]. Fe_3C is known to be less active than the ferrite phase. Therefore, there is a preferential dissolution of ferrite over cementite, working the former as the anode and latter as the cathode, favoring the hydrogen evolved reaction (HER) during the corrosion process [21, 22]. Variation of the corrosion rate for inhibitors II, VI and V at different concentrations are presented in Figure 1. Corrosion parameters were calculated on the basis of LPR corrosion rate test. The inhibition efficiency (η %) and surface coverage (θ) were calculated according to the following equations: where CR_0 is the corrosion rate without inhibitor and CR_i is the corrosion rate when inhibitor is present. It can be seen that the presence of inhibitors		
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results a high decrease in the rate of corrosion. In the case of these inhibitors, the corrosion rate decreases as the inhibitor concentration increases, getting maximum inhibition efficiency ranged between 97.00 and 99.90 % at 100 ppm after 20 hour of exposure (Table 3). This trend may results from the fact that adsorption and surface coverage increase with the increase in concentration; thus the surface is effectively separated from the medium [23]. Table 3 shows the calculated values of corrosion rates, the inhibition efficiencies and the surface coverage in the absence and presence of different concentrations of different inhibitors at 50 °C. The data exhibited that, the corrosion rates, the inhibition efficiencies and the surface coverage are found to depend on the concentrations of the inhibitors. The corrosion rate (CR) are decreased, and the inhibition efficiencies ($\eta\%$) and the surface		
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coverage (θ) are increased with the increase of the surfactant concentrations. This indicates that the inhibitory action of the inhibitors against carbon steel corrosion can be attributed to the adsorption of these molecules on the metal surface, limits the dissolution of carbon steel, and the adsorption amounts of surfactants on		
shows the variation of the corrosion rate with time for carbon steel in CO₂-saturated brine containing 100 ppm from different surfactants at 50 °C. This plot indicates that, the presence of different inhibitors decreases the rate of corrosion. However, the maximum decrease in the corrosion rate was observed for inhibitor (V) and the inhibition efficiency of the investigated surfactants was increased in the following order: V>IV>II>VI>III>I. The highest inhibition efficiency of inhibitors IV and V may be attributed to the following reasons: The higher electron density on the	http://pubs.sciepub.com/jmpc/1/2/2	Kindly take the reference of the source inside your paper to avoid plagiarism complain.

functional groups leads to easier bond formation, greater adsorption, and consequently, higher inhibition		
shows the change in corrosion potential (E_{corr}) with time for carbon steel in CO ₂ -saturated brine in the absence and presence of inhibitor (V). It can be observed, an anodic shift in E_{corr} that is characteristic of anodic and anodic/cathodic inhibitor [28]. It was explained that this shift in E_{corr} is due to active sites blocking effect that occurs when an inhibitor is added [29]. In the case of CO ₂ corrosion the anodic a cathodic reactions are the oxidation of iron and the reduction of hydrogen, respectively [30]. If it is considered that the active sites on the metal surface are the same for both reactions before adding the inhibitor, it is logical the change in E_{corr} when the inhibitor is present because its adsorption change those active sites and therefore the anodic and cathodic reaction rates. Another important aspect is that the change in	http://www.electrochemsci.org/papers/vol5/5060797.pdf	Kindly take the reference of the source inside your paper to avoid plagiarism complain.

Ecorr is proportional to the inhibitor efficiency. Higher efficiencies were observed when the change in Ecorr was toward more noble		
the above results and discussion, the following conclusions are drawn: Novel surfactants in the series of in the series of fatty acids were synthesized based on palm oil and diethanolamine. The corrosion inhibition efficiencies of the synthesized surfactants were studied by linear polarization corrosion rate and Potentiodynamic polarization measurements, and they gave similar results. Test results show that the surfactants studied are efficient inhibitors for	http://pubs.sciepub.com/jmpc/1/2/2	Kindly take the reference of the source inside your paper to avoid plagiarism complain.